Chemical Reactions-Multiple Choice Review

PSI Chemistry

Name_______________________________

1) What are the missing coefficients for the skeleton equation below?
   \[ \text{Al}_2(\text{SO}_4)_{3(aq)} + \text{KOH(aq)} \rightarrow \text{Al(OH)}_{3(aq)} + \text{K}_2\text{SO}_4(aq) \]
   A) 1,3,2,3  
   B) 2,12,4,6  
   C) 4,6,2,3  
   D) 1,6,2,3  
   E) 2,3,1,1

2) What are the missing coefficients for the skeleton equation below?
   \[ \text{Cr(s)} + \text{Fe(NO}_3)_{2(aq)} \rightarrow \text{Fe(s)} + \text{Cr(NO}_3)_{3(aq)} \]
   A) 4,6,6,2  
   B) 2,3,2,3  
   C) 2,3,3,2  
   D) 1,3,3,1  
   E) 2,3,1,2

3) What are the missing coefficients for the skeleton equation below?
   \[ \text{NH}_3(g) + \text{O}_2(g) \rightarrow \text{N}_2(g) + \text{H}_2\text{O(l)} \]
   A) 4,3,2,6  
   B) 2,1,2,3  
   C) 1,3,1,3  
   D) 2,3,2,3  
   E) 3,4,6,2

4) If you rewrite the following word equation as a balanced chemical equation, what will the coefficient and symbol for iodine be?
   \[ \text{bromine + potassium iodide} \rightarrow \text{potassium bromide + iodine} \]
   A) 2I^{-}  
   B) I  
   C) 2I  
   D) I_{2}  
   E) 2I_{2}

5) If you rewrite the following word equation as a balanced chemical equation, what will the coefficient and symbol for fluorine be?
   \[ \text{nitrogen trifluoride} \rightarrow \text{nitrogen + fluorine} \]
   A) 3F  
   B) 6F_{2}  
   C) F_{3}  
   D) 6F  
   E) 3F_{2}
6) What are the missing coefficients for the skeleton equation below?
\[ \text{AlCl}_3 + \text{NaOH} \rightarrow \text{Al(OH)}_3 + \text{NaCl} \]
A) 1,3,1,3
B) 3,1,3,1
C) 1,1,1,3
D) 1,3,3,1
E) 3,1,1,1

7) What are the missing coefficients for the skeleton equation below?
\[ \text{N}_2 + \text{H}_2 \rightarrow \text{NH}_3 \]
A) 1,1,2
B) 1,3,3
C) 3,1,2
D) 1,3,2
E) 2,6,6

8) Aluminum chloride and bubbles of hydrogen gas are produced when metallic aluminum is placed in hydrochloric acid. What is the balanced equation for this reaction?
A) \( \text{H} + \text{AlCl} \rightarrow \text{Al} + \text{HCl} \)
B) \( 2\text{Al} + 6\text{HCl} \rightarrow 2\text{AlCl}_3 + 3\text{H}_2 \)
C) \( \text{Al} + \text{HCl}_3 \rightarrow \text{AlCl}_3 + \text{H} \)
D) \( \text{Al} + 2\text{HCl} \rightarrow \text{AlCl}_2 + \text{H}_2 \)
E) \( \text{H}_2 + \text{AlCl}_3 \rightarrow \text{Al} + 2\text{HCl} \)

9) What does the symbol \( \Delta \) in a chemical equation mean?
A) heat is supplied to the reaction
B) a catalyst is needed
C) yields
D) precipitate

10) When the equation, \( \text{Fe} + \text{Cl}_2 \rightarrow \text{FeCl}_3 \), is balanced, what is the coefficient for \( \text{Cl}_2 \)?
A) 1
B) 2
C) 3
D) 4

11) When the following equation is balanced, what is the coefficient for \( \text{HCl} \)?
\[ \text{Mg(s)} + \text{HCl(aq)} \rightarrow \text{MgCl}_2 (\text{aq}) + \text{H}_2(\text{g}) \]
A) 6
B) 3
C) 1
D) 2
12) Chemical reactions _____.

A) occur only in living organisms  
B) create and destroy atoms  
C) only occur outside living organisms  
D) produce new substances

13) Which of the following is NOT a true statement concerning what happens in all chemical reactions?

A) The ways in which atoms are joined together are changed.  
B) New atoms are formed as products.  
C) The starting materials are named reactants.  
D) The bonds of the reactants are broken and new bonds of the products are formed.  
E) In a word equation representing a chemical reaction, the reactants are written on the left and the products on the right.

14) Chemical equations _____.

A) describe chemical reactions  
B) show how to write chemical formulas  
C) give directions for naming chemical compounds  
D) describe only biological changes

15) A skeleton equation does NOT show which of the following?

A) the correct formulas of the reactants and products  
B) the reactants on the left, the products on the right  
C) an arrow connecting the reactants to the products  
D) the physical states of the substances  
E) the relative amounts of reactants and products

16) Chemical equations describe ______.

A) nuclear reactions  
B) electrochemical processes  
C) chemical reactions  
D) biological reactions  
E) all the above

17) Chemical equations must be balanced to satisfy the ______.

A) law of definite proportions  
B) law of multiple proportions  
C) law of conservation of mass  
D) principle of Avogadro
18) Symbols used in equations, together with the explanations of the symbols, are shown below. Which set is correct?

A) (g), grams  
B) (l), liters  
C) (aq), dissolved in water  
D) (sp), solid product  
E) (lq), liquid  

19) In the chemical equation, \( \text{H}_2\text{O}_2(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g}) \), the \( \text{H}_2\text{O}_2 \) is a _____.

A) product  
B) reactant  
C) catalyst  
D) solid  
E) gas  

20) A catalyst is _____.

A) the product of a reaction  
B) is a reactant  
C) one of the reactants in single-replacement reactions  
D) a solid product of a reaction  
E) a chemical that speeds up the reaction  

21) When the following equation is balanced, \( \text{KClO}_3(\text{s}) \rightarrow \text{KCl}(\text{s}) + \text{O}_2(\text{g}) \), the coefficient of \( \text{KClO}_3 \) is _____.

A) 1  
B) 2  
C) 3  
D) 4  
E) 6  

22) Which of the following is the correct skeleton equation for the reaction that takes place when solid phosphorus combines with oxygen gas to form diphosphorus pentoxide?

A) \( \text{P}(\text{s}) + \text{O}_2(\text{g}) \rightarrow \text{PO}_2(\text{g}) \)  
B) \( \text{P}(\text{s}) + \text{O}(\text{g}) \rightarrow \text{P}_2\text{O}_5(\text{g}) \)  
C) \( \text{P}(\text{s}) + \text{O}_2(\text{g}) \rightarrow \text{P}_2\text{O}_5(\text{g}) \)  
D) \( \text{P}_2\text{O}_5 \rightarrow \text{P}_2(\text{s}) + \text{O}_2(\text{g}) \)  
E) \( \text{P}_2(\text{s}) + \text{O}_5(\text{g}) \rightarrow \text{P}_2\text{O}_5(\text{g}) \)  

23) In every balanced chemical equation, each side of the equation has the same number of _____.

A) atoms  
B) molecules  
C) moles
24) When potassium hydroxide and barium chloride react, potassium chloride and barium hydroxide are formed. The balanced equation for this reaction is _____.

A) KH + BaCl \rightarrow KCl + BaH  
B) KOH + BaCl \rightarrow KCl + BaOH  
C) 2KOH + BaCl_2 \rightarrow 2KCl + Ba(OH)_2  
D) KOH + BaCl_2 \rightarrow K+ BaOH  
E) 2KOH + 2BaCl_2 \rightarrow 2KCl + 2Ba(OH)_2

25) The double arrow symbol indicates _____.

A) that heat must be applied  
B) an incomplete combustion reaction  
C) that a gas is formed by the reaction  
D) that the reaction is reversible

26) If a combination reaction takes place between potassium and chlorine, what is the product?

A) KCl  
B) KCl_2  
C) K_2Cl  
D) PCl  
E) PCl_2

27) The product of a combination reaction is Ba(OH)_2. If one reactant was H_2O what was the other reactant?

A) Ba_2O  
B) BaO  
C) BaH  
D) BaO_2  
E) Ba_2O_7

28) Write a balanced equation for the combination reaction that takes place when iron(III) oxide is formed from its constituent elements.

A) Fe_2 + O_3 \rightarrow Fe_2O_3  
B) 2Fe + 3O \rightarrow Fe_2O_3  
C) 4Fe + 3O_2 \rightarrow 2Fe_2O_3  
D) 3Fe + O \rightarrow Fe_3O  
E) Fe + O_3 \rightarrow FeO_3
29) The reaction, \(2\text{Fe} + 3\text{Cl}_2 \rightarrow 2\text{FeCl}_3\), is an example of which type of reaction?

A) REDOX - combustion reaction  
B) REDOX - disproportionation reaction  
C) REDOX - combination reaction  
D) Precipitation Reaction  
E) Acid/Base Reaction

30) Write a balanced equation to represent the decomposition of lead(IV) oxide into its constituent elements.

A) \(\text{PbO}_2 \rightarrow \text{Pb} + 2\text{O}\)  
B) \(\text{PbO}_2 \rightarrow \text{Pb} + \text{O}_2\)  
C) \(\text{Pb}_2\text{O} \rightarrow 2\text{Pb} + \text{O}\)  
D) \(\text{PbO} \rightarrow \text{Pb} + \text{O}_2\)  
E) \(2\text{PbO} \rightarrow 2\text{Pb} + \text{O}_2\)

31) What is the balanced equation for the reaction that takes place between bromine and sodium iodide?

A) \(\text{Br}_2 + \text{NaI} \rightarrow \text{NaBr}_2 + \text{I}\)  
B) \(\text{Br}_2 + 2\text{NaI} \rightarrow 2\text{NaBr} + \text{I}_2\)  
C) \(\text{Br}_2 + 2\text{NaI} \rightarrow 2\text{NaBr} + 2\text{I}^-\)  
D) \(\text{Br} + \text{NaI}_2 \rightarrow \text{NaBrI}_2\)  
E) \(\text{Br} + \text{NaI}_2 \rightarrow \text{NaBr} + \text{I}_2\)

32) The equation \(\text{Mg(s)} + 2\text{HCl(aq)} \rightarrow \text{MgCl}_2(\text{aq}) + \text{H}_2\) is an example of which type of reaction?

A) REDOX reaction  
B) Acid/Base reaction  
C) REDOX - disproportionation reaction  
D) Precipitation reaction

33) What are the correct formulas and coefficients for the products of this acid/base neutralization reaction?

\(\text{RbOH} + \text{H}_3\text{PO}_4 \rightarrow\)

A) \(\text{Rb}(\text{PO}_4)_3 + \text{H}_2\text{O}\)  
B) \(\text{RbPO}_4 + 2\text{H}_2\text{O}\)  
C) \(\text{Rb}_3\text{PO}_4 + 3\text{H}_2\text{O}\)  
D) \(\text{H}_3\text{Rb} + \text{PO}_4\text{OH}\)  
E) \(3\text{RbH} + \text{H}_2\text{OPO}_4\)

34) The equation \(\text{H}_3\text{PO}_4 + 3\text{KOH} \rightarrow \text{K}_3\text{PO}_4 + 3\text{H}_2\text{O}\) is an example of which type of reaction?

A) REDOX - combination reaction  
B) REDOX - disproportionation reaction
When the equation for the complete combustion of ethanol, C₂H₅OH, is balanced, what is the coefficient for oxygen?

- A) 1
- B) 3
- C) 6
- D) 7
- E) 14

The equation 2C₃H₇OH + 9O₂ → 6CO₂ + 8H₂O is an example of which type of reaction?

- A) REDOX - combustion
- B) REDOX - synthesis
- C) Acid/base reaction
- D) Precipitation reaction

Which of the following is NOT true concerning the decomposition of a simple binary compound?

- A) The products are unpredictable.
- B) The products are the constituent elements.
- C) The reactant is a single substance.
- D) The reactant could be an ionic or a molecular compound.
- E) Energy is usually required.

Which of the following is TRUE regarding a precipitation reaction.

- A) Both products must be soluble in water
- B) At least one of the products will be insoluble in water
- C) Spectator ions are always included in the reaction
- D) Electrons are gained and lost
- E) H⁺ ions are transferred from one molecule to another

Which of the following ions will NEVER form a precipitate?

- A) K⁺
- B) Ca²⁺
- C) CO₃²⁻
- D) SO₄²⁻
- E) Pb²⁺

Combining aqueous solutions of BaI₂ and Na₂SO₄ affords a precipitate of BaSO₄. Which ion(s) is/are spectator ions in the reaction?
41) Which ion(s) is/are spectator ions in the formation of a precipitate of AgCl via combining aqueous solutions of CoCl₂ and AgNO₃?

A) Co²⁺ and NO₃⁻
B) NO₃⁻ and Cl⁻
C) Co²⁺ and Ag⁺
D) Cl⁻
E) NO₃⁻

42) The balanced net ionic equation for precipitation of CaCO₃ when aqueous solutions of Na₂CO₃ and CaCl₂ are mixed is ____________.

A) 2Na⁺(aq) + CO₃²⁻(aq) → Na₂CO₃(aq)
B) 2Na⁺(aq) + 2Cl⁻(aq) → 2NaCl(aq)
C) Na⁺(aq) + Cl⁻(aq) → NaCl(aq)
D) Ca²⁺(aq) + CO₃²⁻(aq) → CaCO₃(s)
E) Na₂CO₃(aq) + CaCl₂(aq) → 2NaCl(aq) + CaCO₃(s)

43) When aqueous solutions of AgNO₃ and KI are mixed, AgI precipitates. The balanced net ionic equation is:

A) Ag⁺(aq) + I⁻(aq) → AgI(s)
B) Ag⁺(aq) + NO₃⁻(aq) → Ag NO₃(s)
C) Ag⁺(aq) + NO₃⁻(aq) → Ag NO₃(aq)
D) Ag NO₃(aq) + KI(aq) → AgI(s) + KNO₃(aq)
E) Ag NO₃(aq) + KI(aq) → AgI(aq) + KNO₃(s)

44) A precipitation reaction takes place when aqueous cobalt(III) chloride reacts with aqueous lithium hydroxide. One of the products of this reaction would be _____.

A) Co(OH)₃ (s)
B) Co(OH)₂ (s)
C) LiCO₃ (s)
D) LiCl₃ (aq)
E) Cl₃OH (s)

45) What is the driving force in the following reaction?

Ni(NO₃)₂(aq) + K₂S(aq) → NiS + 2KNO₃(aq)

A) A gas is formed.
B) A precipitate is formed.
C) Ionic compounds are reactants.
D) Ionic compounds are products.
E) Heat is required.

46) A precipitation reaction takes place when aqueous Na₂CO₃ reacts with aqueous Sn(NO₃)₂. You would expect one of the products of this reaction to be ______.
A) NaNO₃ (aq)  
B) NaSn (s)  
C) Sn(CO₃)₂ (s)  
D) CNO₃

47) Which of the following solutions, when added to a solution of Na₂SO₄ would form a precipitate?  
I. AgNO₃(aq)  
II. KNO₃(aq)  
III. CaCl₂ (aq)  
A) I only  
B) II only  
C) III only  
D) I and II only  
E) I and III only

48) Which of the following solutions, when added to a solution of Sr(NO₃)₂(aq), would form a precipitate?  
I. Na₂SO₄ (aq)  
II. HgC₂H₃O₂(aq)  
III. NH₄Cl  
A) I only  
B) II only  
C) III only  
D) I and II only  
E) I, II, and III

49) Which of the following represents the correct net-ionic equation for the reaction of calcium nitrate with potassium phosphate?  
A) Ca²⁺(aq) + PO₄³⁻ (aq) --> CaPO₄(s)  
B) 2Ca²⁺ (aq) + 3PO₄³⁻ (aq) --> Ca₂(PO₄)₃(s)  
C) 3Ca²⁺ (aq) + 2PO₄³⁻(aq) --> Ca₃(PO₄)₂(s)  
D) K⁺(aq) + NO₃⁻(aq) --> KNO₃(s)  
E) 2NO₃⁻(aq) + Ca²⁺ --> Ca(NO₃)₂(s)

50) Which of the following is/are TRUE regarding oxidation/reduction reactions?  
I. Oxidation involves the gain of electrons  
II. electrons are transferred from one substance to another  
III. The number of electrons lost/gained must be equal  
A. I only  
B. II only  
C. III only  
D. II and III  
E. I, II, and III
51) What would be the correct oxidation state of N in the nitrite ion (NO$_2^-$)?
A. +1
B. +3
C. +5
D. -3
E. -1

52) What would be the correct oxidation state of N and O respectively in KNO$_3$?
A. +1 and -2
B. +3 and -2
C. +5 and -2
D. +5 and -1
E. -3 and -2

53. Which of the following compounds contain a Cl atom with a +7 oxidation state?
A. Cl$_2$
B. HOCl
C. ClO$_3^-$
D. KCl
E. KClO$_4$

54. In the following reaction, 2Al + 6HCl $\rightarrow$ 2AlCl$_3$ + 3H$_2$, the oxidation state for aluminum goes from...
A) 0 --> +1
B) 0 --> +3
C) +3 --> +3
D) +3 --> -3
E) 0 --> -3

55. In which of the following compounds would hydrogen have an oxidation state of zero?
A) HI
B) NaH
C) Ca(OH)$_2$
D) NH$_4$Cl
E) H$_2$

56. Which of the following represents an oxidation?
A. Na$^+$ --> Na
B. N$_2$ --> 2N$^3-$
C. Mn$^{7+}$ --> Mn$^{2+}$
D. Fe --> Fe$^{3+}$
E. None of these are oxidations

57. In which of the following reactions does hydrogen get reduced?
A. 2H$_2$ + O$_2$ $\rightarrow$ 2H$_2$O
B. Cl$_2$ + H$_2$ $\rightarrow$ 2HCl
C. \(2\text{NaH} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + 3\text{H}_2\)
D. \(\text{Mg} + 2\text{H}^+ \rightarrow \text{Mg}^{2+} + \text{H}_2\)
E. H never gets reduced

58. In a combustion reaction, one of the reactants is _____.

A) hydrogen
B) nitrogen
C) oxygen
D) a metal
E) a binary ionic compound

59) The products of a combustion reaction do NOT include _____.

A) water
B) carbon dioxide
C) carbon monoxide
D) heat
E) hydrogen

60) The complete combustion of which of the following substances produces carbon dioxide and water?

A) \(\text{C}_8\text{H}_{18}\)
B) \(\text{K}_2\text{CO}_3\)
C) \(\text{CaHCO}_3\)
D) NO
E) \(\text{H}_2\text{S}\)

61) Which of the following is the correctly balanced equation for the incomplete combustion of heptene, \(\text{C}_7\text{H}_{14}\)?

A) \(\text{C}_7\text{H}_{14} + 14\text{O} \rightarrow 7\text{CO} + 7\text{H}_2\text{O}\)
B) \(\text{C}_7\text{H}_{14} + 7\text{O}_2 \rightarrow 7\text{CO} + 7\text{H}_2\text{O}\)
C) \(2\text{C}_7\text{H}_{14} + 21\text{O}_2 \rightarrow 14\text{CO}_2 + 14\text{H}_2\text{O}\)
D) \(\text{C}_7\text{H}_{14} + \text{O}_2 \rightarrow \text{C}_7\text{O}_2 + 7\text{H}_2\)
E) \(\text{C}_7\text{H}_{14} + 7\text{O}_2 \rightarrow 7\text{CO}_2 + 7\text{H}_2\text{O}\)

62) Which of the following is TRUE regarding the following reaction?

\(\text{F}_2(g) + 2\text{I}^-(aq) \rightarrow \text{I}_2(s) + 2\text{F}^-(aq)\)

I. Flourine gets oxidized
II. Iodide ions lose electrons
III. It is a disproportionation reaction

A) I only
B) II only
63) Which of the following would NOT be an oxidation/reduction reaction?
A. Na(s) + I₂(s) → 2NaI
B. Ca²⁺(aq) + CO₃²⁻(aq) → CaCO₃(s)
C. 2H₂O(s) → 2H₂(g) + O₂(g)
D. Fe²⁺(aq) + H⁺(aq) + CrO₄²⁻(aq) → Fe³⁺(aq) + H₂O(l) + Cr³⁺(aq)
E. 2Fe₂O₃(s) → 4Fe(s) + 3O₂(g)

64) What would be the products of the acid/base reaction between KOH (aq) and HNO₃ (aq)?
A) K⁺ and H⁺
B) H⁺ and OH⁻
C) K⁺ and NO₃⁻
D) H₂O + K⁺ + NO₃⁻
E) OH⁻ only

65) The net ionic equation for the reaction between aqueous solutions of HF and KOH is

A) HF + KOH → H₂O + K⁺ + F⁻
B) HF + OH⁻ → H₂O + F⁻
C) HF + K⁺ + OH⁻ → H₂O + KF
D) H⁺ + OH⁻ → H₂O
E) H⁺ + F⁻ + K⁺ + OH⁻ → H₂O + K⁺ + F⁻

66) When H₂SO₄ is neutralized by NaOH in aqueous solution, the net ionic equation is

A) SO₄²⁻(aq) + 2Na⁺(aq) → Na₂SO₄(aq)
B) SO₄²⁻(aq) + 2Na⁺(aq) → Na₂SO₄(s)
C) H⁺(aq) + OH⁻(aq) → H₂O(l)
D) H₂SO₄(aq) + 2OH⁻(aq) → 2H₂O(l) + SO₄²⁻(aq)
E) 2H⁺(aq) + 2NaOH(aq) → 2H₂O(l) + 2Na⁺(aq)

67) Which of the following would be TRUE regarding acid/base reactions?
I. Acids donate electrons
II. Bases accept H⁺ ions
III. In a neutralization reaction, H₂O is a product

A) I only
B) II only
C) III only
D) I and III
E) II and III
68) Which of the following would be a disproportionation reaction?

A) \( \text{Fe}^{2+}(aq) \rightarrow \text{Fe}^{3+}(aq) + \text{Fe}(s) \)

B) \( 2\text{Cl}^{-}(aq) + \text{Pb}^{2+}(aq) \rightarrow \text{PbCl}_2(s) \)

C) \( \text{PCl}_5(g) \rightarrow \text{PCl}_3(g) + \text{Cl}_2(g) \)

D) \( \text{HCN}(aq) + \text{OH}^{-}(aq) \rightarrow \text{H}_2\text{O}(l) + \text{CN}^{-}(aq) \)

E) None of these

For the following, indicate if the first statement is true or false, if the second statement is true or false, and then if the second statement is an explanation of the first. These questions are designed to mimic a type of question found on the SAT II chemistry exam.

69. Acids donate H+ ions \hspace{1cm} \text{BECAUSE} \hspace{1cm} \text{Bases accept H+ ions}

70. Spectators ions are not included in a reaction \hspace{1cm} \text{BECAUSE} \hspace{1cm} \text{spectator ions are typically insoluble}

71. Iron gets oxidized in the reaction \hspace{1cm} \text{BECAUSE} \hspace{1cm} \text{Fe lost electrons}

\( \text{Fe}(s) \rightarrow \text{Fe}^{3+}(aq) \)

Answers:
1) D  
2) C  
3) A  
4) D  
5) E  
6) A  
7) D  
8) B  
9) A  
10) C  
11) D  
12) D  
13) B  
14) A  
15) E  
16) E  
17) C  
18) C  
19) B  
20) E  
21) B  
22) C  
23) A  
24) C  
25) D  
26) A  
27) B  
28) C  
29) C  
30) B  
31) B  
32) A  
33) C  
34) C  
35) D  
36) A  
37) A  
38) B  
39) A  
40) D  
41) A  
42) D  
43) A  
44) A  
45) B  
46) A  
47) E  
48) A  
49) C  
50) D  
51) B  
52) C  
53) E  
54) B  
55) E  
56) D  
57) D  
58) C  
59) E  
60) A  
61) B  
62) B
63) B
64) D
65) B
66) C
67) E
68) A
69) T,T,
   No
70) T,F,
   No
71) T,T,Yes